

Supplementary material for the article:

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Interaction of dinuclear cadmium(II) 5–Cl–salicylaldehyde complexes with calf–thymus DNA

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Supplementary Information

S1. Interaction with CT DNA

The binding constant, K_b , can be obtained by monitoring the changes in the absorbance at the corresponding λ_{\max} with increasing concentrations of CT DNA and it is given by the ratio of slope to the y intercept in plots $\frac{[\text{DNA}]}{(\varepsilon_A - \varepsilon_f)}$ versus [DNA], according to the Wolfe–Shimer equation

[S1]:

$$\frac{[\text{DNA}]}{(\varepsilon_A - \varepsilon_f)} = \frac{[\text{DNA}]}{(\varepsilon_b - \varepsilon_f)} + \frac{1}{K_b(\varepsilon_b - \varepsilon_f)} \quad (\text{eq. S1})$$

where [DNA] is the concentration of DNA in base pairs, $\varepsilon_A = A_{\text{obsd}}/[\text{compound}]$, ε_f = the extinction coefficient for the free compound and ε_b = the extinction coefficient for the compound in the fully bound form.

S2. Competitive studies with EB

The Stern–Volmer constant K_{SV} is used to evaluate the quenching efficiency for each compound according to the Stern–Volmer equation [S2]:

$$\frac{I_0}{I} = 1 + K_{SV}[Q] \quad (\text{eq. S2})$$

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where I_0 and I are the emission intensities in the absence and the presence of the quencher, respectively, $[Q]$ is the concentration of the quencher (i.e. complexes **1–5**); K_{SV} is obtained from the Stern–Volmer plots by the slope of the diagram $\frac{I_0}{I}$ vs $[Q]$.

References

- [S1] A. Wolfe, G. Shimer, T. Meehan, *Biochemistry* 26 (1987) 6392–6396.
[S2] J.R. Lakowicz, *Principles of Fluorescence Spectroscopy*, 3rd ed. Plenum Press, New York, (2006).

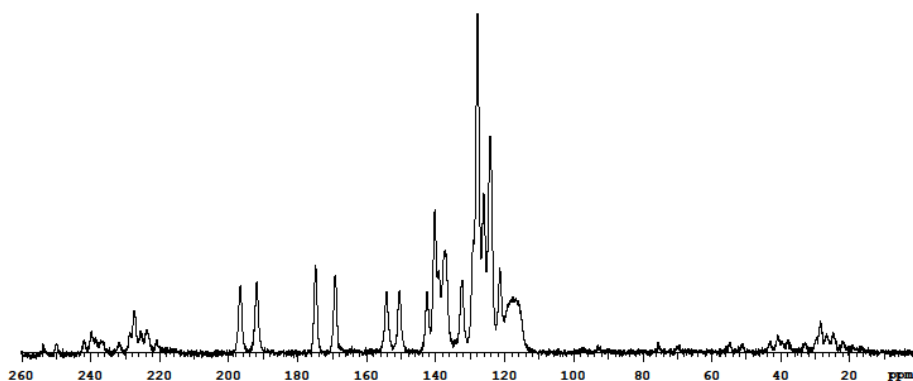


Fig. S1. ^{13}C -NMR spectra in solid state of $[\text{Cd}(5\text{-Cl-salo})_2(\text{phen})]_2$, (**3**)

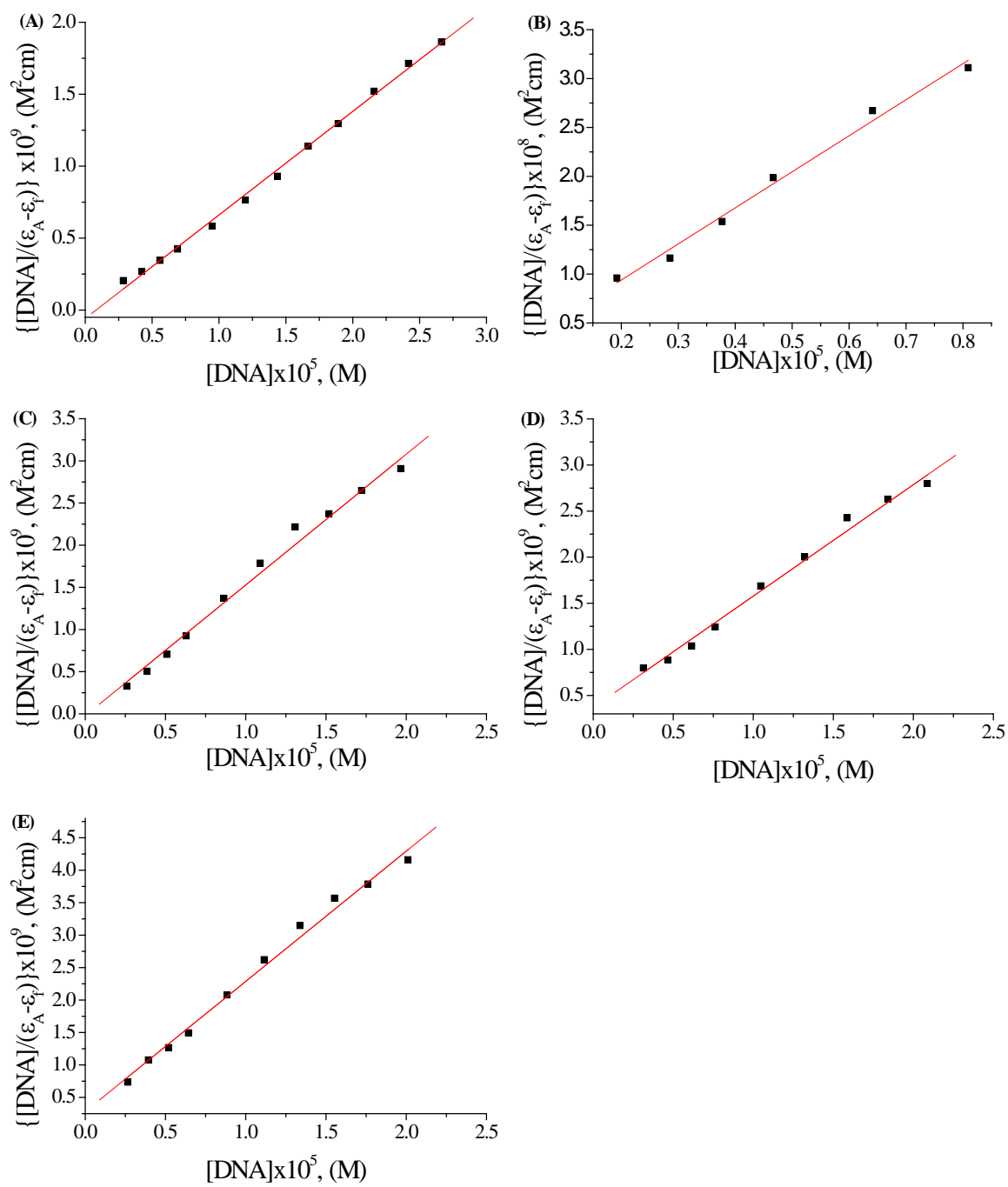


Fig. S2. (A)–(E) Plot of $\frac{[DNA]}{(\epsilon_A - \epsilon_f)}$ vs $[DNA]$ for complexes 1–5, respectively.

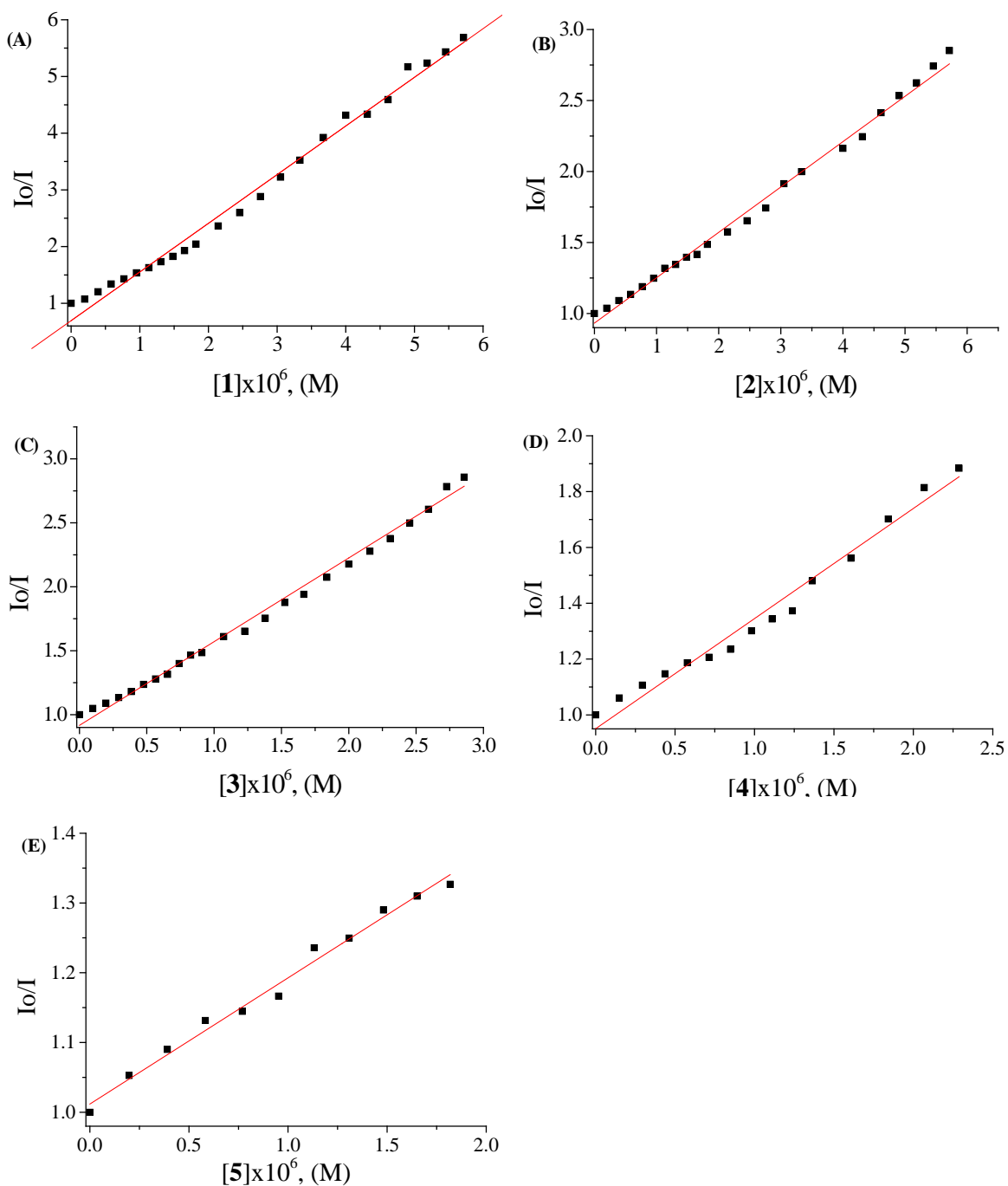


Fig. S3. (A)–(E) Stern–Volmer quenching plot of EB bound to CT DNA for complexes 1–5, respectively.